Paper: __Chemistry_____

Month Test: ___5th Term_____

Theme/Unit: _____

Objective/Subjective:

Name: _____

ID: _____

class:__9th_____Section: ___

Total Marks: __55_____

Obt. Marks: _____

Weekly Test: _____/____

Worksheets: _____/_____

Grand Total: _____

Q: no: 1: Encircle the correct option. /10

1) Which one of the following non-metal is lustrous?

   a) Sulphur   b) Phosphorus   c) Iodine   d) Carbon

2) Metal can form ions carrying charges:

   a) Uni-positive   b) Di-positive   c) Tri-positive   d) All of them

3) Which is the yellow soft metal?

   a) Silver   b) Platinum   c) Gold   d) none of these

4) The mass of one molecule of water is:

   a) 18 amu   b) 18 g   c) 18 mg   d) 18 kg

5) One amu (atomic mass unit) is equal to:

   a)1.66×10^{-24}mg   b) 1.66×10^{-24} g   c)1.66×10^{-24}kg   d)1.66×10^{-23} g

6) Oxidation is:

   a) Removal of Oxygen   b) Removal of Hydrogen   c) Gaining of Electrons

   d) None of these

7) The oxidation number of chromium in $K_2Cr_2O_7$ is:

   a) +2   b) +6   c) +7   d) +14

8) The formula of rust is:

   a) $Fe_2O_3 . nH_2O$ b) Fe_2O_3 c) $Fe(OH)_3 . nH_2O$ d) $Fe(OH)_3$

9) Sodium is extremely reactive metal, but it does not react with:

    a) Hydrogen   b) Nitrogen   c) Sulphur   d) Phosphorus

10) Valency of Sodium is:

    a) 1   b) 2   c) 2, 3   d) 4

Q: no: 2: Short answer questions.   **/12*2=24**

1) Write two difference between Ions and Free Radicals.

2) Define empirical formula with an example.

3) Define atomic mass unit. Why is it needed?

4) Why is an iron grill painted frequently?

5) What is the nature of electrode used in electrolyzing of chromium?

6) Where do the electrons flow from Zn electrode in Daniel's cell?

7) Why is O_2 necessary for rusting?

8) Why are silver and gold least reactive?

9) Why is copper used for making electrical wires?

10) How electro positivity depends upon size and nuclear charge of an atom?

11) Define Matter and Substance.

12) Define molecular formula with an example.

Q: no: 3:   / 5+5=10

(a) : Discuss the electrolysis of water.

b) How many atoms are required to prepare 60 g of HNO_3?

Q: no: 4   /6+5=11

a) Give the reaction of sodium with : H_2O, O_2, CL_2 and H_2.

b) Calculate the formula mass of Potassium sulphate K_2SO_4.