

Paper: ChemistryTotal Marks: 12Month Test: Annual

Obt. Marks: _____

Theme/Unit: CompleteGrand Total: 75

Objective: _____

Signature: _____

Time: 15 mins

Roll No: _____

class: 9th

Section: _____

**Q. No. 1: Encircle the correct option:**

- The mass of proton is more than that of electron:
 - 1850 times
 - 1856 times
 - 1840 times
 - 1830 times
- Molecular mass of H_3PO_4 is:
 - 99gmol^{-1}
 - 98gmol^{-1}
 - 89gmol^{-1}
 - None
- Which elements are use in powdered form in fireworks:
 - Magnesium
 - Cobalt
 - Nickl
 - Lithium
- Shielding remain the same in a period if we move from:
 - Right to left
 - Down to the period
 - Left to right
 - None
- Group 17 has electrone in its valence shell?
 - 10
 - 7
 - 17
 - 18
- Identify the compound which is not soluble in water.
 - C_6H_6
 - Nacl
 - KBr
 - O_2
- Charles's law presented in:
 - 1787
 - 1788
 - 1789
 - 1888
- Grease and paints are soluble in:
 - Carbon tetrachloride
 - Carbon tetrafluoride
 - Carbon silicate
 - All
- Tyndall effect is shown by:
 - Sugar solution
 - Paints
 - Jelly
 - Chalk solution
- A substance that oxidizes itself and reduce other called?
 - Oxidizing agent
 - Reducing agent
 - Both
 - None
- The formula of rust is:
 - $Fe_2O_3 \cdot nH_2O$
 - Fe_2O_3
 - $Fe(OH)_3$
 - Fecl
- Which of the following is less malleable?
 - Iron
 - Gold
 - Silver
 - Sodium

Paper: ChemistryTotal Marks: 60Month Test: Annual

Obt. Marks: _____

Theme/Unit: CompleteGrand Total: 75

Subjective:

Signature: _____

Time: _____

Roll No: _____

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Section: _____

**Q. No. 1: Answer the short Questions: /10**

- 1) Why is ionization energy of Na less than mg?
- 2) W are silver and gold least reactive?
- 3) What is the principle of electroplating?
- 4) What is rust? Write down its equation.
- 5) Write down the structure and function of salt bridge.

Q. No. 3: Answer the short Questions: /10

- 1) Why the suspension does not form a homogenous mixture?
- 2) Why is a solution considered mixture?
- 3) Define transition state, with example?
- 4) Define hydrogen bonding?
- 5) Differentiate b/w lone pair and bond pair of electrons?

Q. No. 4: Answer the short Questions: /10

- 1) Define effective nuclear charge?
- 2) Why the elements are called S or P block elements?
- 3) Define electronic configuration?
- 4) What is the relationship between empirical formula and formula unit?
- 5) Difference b/w atom and ion?

Part – II**Q. No. 5: (a).** Describe the differences b/w compound and mixture? /5**(b).** How Rutherford discovered that atoms has nucleus located at the centre of the atom?/6**Q. No. 6: (a).** What is covalent compound? Describe its general properties. /5**(b).** Briefly explain the factors on which evaporation depend? /6**Q. No. 7: (a).** Define solubility? Describe general principle of solubility? /5**(b).** Explain the working of Nelson's cell? /6