Paper: _____Chemistry_____   Total Marks: _____75_____

Month Test: ___May_____   Obt. Marks: _____

Theme/Unit: _____1_____   Grand Total: _____

Objective/Subjective: _____   ID: _____   Time: _____

Name: _____   class: _____9th_____   Section: _____B_____

Q NO. 1:   Choose the right answer.   /12
1) Brass is a mixture of:
   a) Copper and Zinc   b) Zinc and Iron
   c) Copper and Iron   d) Iron and tin
2) Caustic soda is common name of:
   a) Na_2CO_3 b) NaOH   c) NaCl   d) HCl
3) Plasma is made up of:
   a) Cations   b) Anions   c) Both   d) None
4) Concentration of Oxygen in atmosphere is:
   a) 47%   b) 21%   c) 86%   d) 16%
5) Empirical formula of Benzene is:
   a) C_6H_6 b) CH_2O c) CH   d) HO
6) 1 amu is equal to:
   a) 1.66×10^{-24} g   b) 1.66×10^{24} g   c) 1.66×10^{23} g   d) 1.66×10^{-23} g
7) Majority of elements exist in the form of:
   a) Plasma   b) Solids   c) Liquids   d) Gases
8) The valency of Aluminium is :
   a) 1   b) 2   c) 3   d) 4
9) 98g of Sulphuric acid is equal to:
   a) 98 mole   b) 100 mole   c) 1 mole   d) 10 mole
10)   About 80% of elements are:
   a) Metalloids   b) metals   c) non-metals   d) Transition elements
11)   The element symbol of Silver is:
   a) S   b) Si   c) Hg   d) Ag
12)   Mass of neutron is:
   a) 1.0087 amu   b) 1.0073 amu   c) 1.009 amu   d) 1.0075 amu

Q No. 2: Short Questions.   /30
1) What are the applications of nuclear chemistry?
2) Differentiate between organic and inorganic chemistry.
3) Define Valency.
4) What is a free radical? How it can be formed?
5) What is Avogadro's number?
6) Briefly explain the concept of mole.
7) Differentiate between atomic number and mass number.
8) How the development of science has made our life easy?
9) Differentiate between molecule and molecular ion?
10)   Define Biochemistry.
11)   Define gram formula mass with example.
12)   What is meant by corpuscular nature of matter?
13)   Differentiate between qualitative and quantitative analysis.
14)   Define empirical formula.
15)   What is relative atomic mass?

Q No. 3: Give detailed answers.   /33
1) a) Write different types of molecule.   /6
     b) Calculate the formula mass of Potassium Sulphate K_2SO_4.   /5
2) a) Give steps to write a chemical formula.   /6
     b) Calculate the protons and neutrons of an atom having A=238 and Z=92.   /5
3) a) Draw summary of molar calculations?   /6
     b) There are 3.01×10^{23} molecules of CO_2 are present in a container. Calculate its number of moles and mass.   /5