

Paper: ChemistryTotal Marks: 75Month Test: May

Obt. Marks: _____

Theme/Unit: 1

Grand Total: _____

Objective/Subjective:

ID: _____

Time: _____

Name: _____

class: 9thSection: B

Q NO. 1: Choose the right answer. /12

1) Brass is a mixture of:

- a) Copper and Zinc b) Zinc and Iron
c) Copper and Iron d) Iron and tin

2) Caustic soda is common name of:

- a) Na_2CO_3 b) NaOH c) NaCl d) HCl

3) Plasma is made up of:

- a) Cations b) Anions c) Both d) None

4) Concentration of Oxygen in atmosphere is:

- a) 47% b) 21% c) 86% d) 16%

5) Empirical formula of Benzene is:

- a) C_6H_6 b) CH_2O c) CH d) HO

6) 1 amu is equal to:

- a) 1.66×10^{-24} g b) 1.66×10^{24} g c) 1.66×10^{-23} g d) 1.66×10^{-23} g

7) Majority of elements exist in the form of:

- a) Plasma b) Solids c) Liquids d) Gases

8) The valency of Aluminium is :

- a) 1 b) 2 c) 3 d) 4

9) 98g of Sulphuric acid is equal to:

- a) 98 mole b) 100 mole c) 1 mole d) 10 mole

10) About 80% of elements are:

- a) Metalloids b) metals c) non-metals d) Transition elements

11) The element symbol of Silver is:

- a) S b) Si c) Hg d) Ag

12) Mass of neutron is:

- a) 1.0087 amu b) 1.0073 amu c) 1.009 amu d) 1.0075 amu

Q No. 2: Short Questions. /30

- 1) What are the applications of nuclear chemistry?
- 2) Differentiate between organic and inorganic chemistry.
- 3) Define Valency.
- 4) What is a free radical? How it can be formed?
- 5) What is Avogadro's number?
- 6) Briefly explain the concept of mole.
- 7) Differentiate between atomic number and mass number.
- 8) How the development of science has made our life easy?
- 9) Differentiate between molecule and molecular ion?
- 10) Define Biochemistry.
- 11) Define gram formula mass with example.
- 12) What is meant by corpuscular nature of matter?
- 13) Differentiate between qualitative and quantitative analysis.
- 14) Define empirical formula.
- 15) What is relative atomic mass?

Q No. 3: Give detailed answers. /33

- 1) a) Write different types of molecule. /6
b) Calculate the formula mass of Potassium Sulphate K_2SO_4 . /5
- 2) a) Give steps to write a chemical formula. /6
b) Calculate the protons and neutrons of an atom having $A=238$ and $Z=92$. /5
- 3) a) Draw summary of molar calculations? /6
b) There are 3.01×10^{23} molecules of CO_2 are present in a container. Calculate its number of moles and mass. /5